r  **NEMETALI, KISELINE, LUŽINE I SOLI - primjer IZ , 8.razred**

1. Razvrstaj navedene spojeve: KOH, NaOH, HNO3, CO, AlCl3, H2CO3, SO2, Mg(OH)2 , CuSO4, H3PO4, H2S, LiOH, NaCl,

 a) kiseline \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ b) hidroksidi \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

 c) oksidi \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ d) soli\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ 4

1. Prikaži kemijske reakcije: a) otapanje sumporova dioksida u vodi\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ 2

 b) reakciju gašenja živog vapna\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ 2

 c) nastajanje magnezijeva oksida \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ 2

3 . Napiši kemijske **nazive** spojeva:

 a) H2SO3\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ b) NaOH \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

 c) Mg(OH)2 \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ d) H2CO3  \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ 4

 4. Popuni tablicu podacima koji nedostaju:

|  |  |  |  |
| --- | --- | --- | --- |
| indikator | boja indikatora u kiselini | indikator | boja indikatora u lužini |
| \_\_\_\_\_\_\_\_\_\_ lakmus |  | \_\_\_\_\_\_\_\_\_\_\_ lakmus |  |
| univerzalni |  | univerzalni |  |
| metiloranž |  |  | crveno-ljubičasta |
| sok crvenog kupusa |  | sok crvenog kupusa |  |

 4

5. Što su lužine?\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ 1

7. Neutralizacija je\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ 1

8. Vodenim otopinama daju kisela svojstva(zaokruži točan odgovor) : a) H3O + ioni b) OH - ioni 1

 a) navedi naziv tih iona\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ 1

9. Prikaži ionizaciju spojeva u vodi: a) H2SO4\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

 b)NaOH\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ 2

10. pH neke otopine je 2. Kakva je ta otopina? \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ 2

11. Zaokruži slovo ispred točnih tvrdnji:

 Željezo je postojano: a) na granici voda- zrak

 b) na suhom zraku

 c) u dodiru s manje plemenitim metalom (Zn)

 d)u dodiru s plemenitijim metalom (Cu) 2

 12. Napiši formule kemijskih spojeva:

 a) kalcijev oksid: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ b) sumporov (IV)oksid\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

 c) karbonatna kiselina: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ d)aluminijev klorid: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ 4

 13. Navedi 3 svojstva natrija: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ 2

 14. Zašto je koncentrirana sumporna kiselina jako dehidratacijsko sredstvo?\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

 \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ 1

16. Otopinu kiseline razrijeđujemo pažljivim dodavanjem \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ u \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_, a nikako \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ u \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ 1

17. Izgrađujem školjke,mramor, planinske masive,puževe kućice. Nisam topljiv u vodi. Moje je ime\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_,

 a kemijska formula\_\_\_\_\_\_\_\_\_\_\_\_\_ . 2

18.Što su hidratne soli?\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ 1

19. Napiši sljedeće ione: a) sulfidni ion\_\_\_\_\_\_\_\_\_ b) nitratni ion\_\_\_\_\_\_\_\_\_ c) fosfatni ion\_\_\_\_\_\_\_\_\_ 3

20. Koji se navedenih hidroksida najbolje otapa u vodi?

 a) kalcijev hidroksid b) natrijev hidroksid c) željezov(III) hidroksid d) željezov(II) hidroksid 1

21. Niz kemijskih promjena opisanih pod a) i b) prikaži kemijskim jednadžbama. 4

 a) Ca CaO Ca(OH)2  b) S SO2  H2SO3

 \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

 \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

22. Što omogućuje električnu vodljivost vodenih otopina kiselina , lužina i soli?\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ 1

23. Dovrši kemijske reakcije:

 a) Na + Cl2 → \_\_\_\_\_\_\_\_\_\_\_ 2

 b) Mg + HCl → \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ 2

 c) KOH + HNO3 → \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ 2

24. Napiši formule spojeva: 3

 a) željezov (III) klorid \_\_\_\_\_\_\_\_\_ b) kalcijev karbonat \_\_\_\_\_\_\_\_\_\_ c) natrijev nitrat \_\_\_\_\_\_\_\_\_

25. Napiši imena soli: a) K2SO4\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ b) ZnCl2 \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ c) Fe(NO3)2\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ 3

26. Što su karbonati? \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ 1

27. Kemijskom jednadžbom prikaži pirolizu modre galice i imenuj nastale produkte. 1

 \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ + \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

 Produkti u reakciji jesu: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ 2

 Želim ti puno uspjeha !

r

r B– grupa Učenik/ca: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ Razred: \_\_\_\_ Bodovi/ocjena :\_\_\_\_\_\_\_\_

1. Razvrstaj navedene spojeve: NaOH, LiOH, HCl, CO, AlBr3, H2SO3, SO3, Mg(OH)2 , CuSO4, H3PO4, H2S, KOH, CuCl2

 a) kiseline \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ b) hidroksidi \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

 c) oksidi \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ d) soli\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ 4

1. Prikaži kemijske reakcije: a) otapanje sumporova dioksida u vodi\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ 2

 b) reakciju nastajanja gašenog vapna\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ 2

 c) nastajanje kalcijeva oksida \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ 2

3 . Napiši kemijske nazivespojeva:

 a) H2CO3\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ b) NaOH \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

 c) Ca(OH)2 \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ d) H2SO3  \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ 4

 4. Popuni tablicu podacima koji nedostaju:

|  |  |  |  |
| --- | --- | --- | --- |
| indikator | boja indikatora u kiselini | indikator | boja indikatora u lužini |
| \_\_\_\_\_\_\_\_\_\_ lakmus |  | \_\_\_\_\_\_\_\_\_\_\_ lakmus |  |
| univerzalni |  | univerzalni |  |
| metiloranž |  |  | crveno-ljubičasta |
| sok crvenog kupusa |  | sok crvenog kupusa |  |

 4

5. Lužine su\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ 1

7. Što je neutralizacija?\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ 1

8. Vodenim otopinama daju lužnata svojstva(zaokruži točan odgovor) : a) H3O + ioni b) OH - ioni 1

 a) navedi naziv tih iona\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ 1

9. Prikaži ionizaciju spojeva u vodi: a) H2CO3\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

 b) NaOH\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ 2

10. pH neke otopine je 8. Kakva je ta otopina? \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ 2

11. Zaokruži slovo ispred točnih tvrdnji:

 Željezo nije postojano: a) na granici voda- zrak

 b) na suhom zraku

 c) u dodiru s manje plemenitim metalom (Zn)

 d)u dodiru s plemenitijim metalom (Cu) 2

 12. Napiši formule kemijskih spojeva:

 a) magnezijev oksid: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ b) sumporov (IV)oksid\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

 c) karbonatna kiselina: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ d)aluminijev klorid: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ 4

 13. Navedi 3 svojstva natrija: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ 2

 14. Zašto je koncentrirana sumporna kiselina jako dehidratacijsko sredstvo?\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

 \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ 1

16. Otopinu kiseline razrijeđujemo pažljivim dodavanjem \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ u \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_, a nikako \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ u \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ 1

17. Izgrađujem školjke,mramor, planinske masive,puževe kućice. Nisam topljiv u vodi. Moje je ime\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_,

 a kemijska formula\_\_\_\_\_\_\_\_\_\_\_\_\_ . 2

18.Što su hidratne soli?\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ 1

19. Napiši sljedeće ione: a) sulfidni ion\_\_\_\_\_\_\_\_\_ b) karbonatni ion\_\_\_\_\_\_\_\_\_ c)nitratni ion\_\_\_\_\_\_\_\_\_ 3

20. Koji se navedenih hidroksida najbolje otapa u vodi?

 a) magnezijev hidroksid b) natrijev hidroksid c) željezov(II) hidroksid d) željezov(III) hidroksid 1

21. Niz kemijskih promjena opisanih pod a) i b) prikaži kemijskim jednadžbama. 4

 a) Mg MgO Mg(OH)2  b) C CO2  H2CO3

 \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

 \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

22. Zbog čega vodene otopine kiselina, lužina i soli provode električnu struju?\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ 1

23. Dovrši kemijske reakcije:

 a) Zn + Cl2 → \_\_\_\_\_\_\_\_\_\_\_ 2

 b) Ca + HCl → \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ 2

 c) NaOH + HNO3 → \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ 2

24. Napiši formule spojeva: 3

 a) željezov (II) klorid \_\_\_\_\_\_\_\_\_ b) kalcijev karbonat \_\_\_\_\_\_\_\_\_\_ c) natrijev nitrat \_\_\_\_\_\_\_\_\_

25. Napiši imena soli: a) Na2SO4\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ b) MgCl2 \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ c) Fe(NO3)3\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ 3

26. Što su sulfati? \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ 1

27. Kemijskom jednadžbom prikaži pirolizu modre galice i imenuj nastale produkte. 1

 \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ + \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

 Produkti u reakciji jesu: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ 2

 Želim ti puno uspjeha !